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Designing A Hand

Warmer

of some chemical reactions. One example is the hand warmer.

There are several types of hand warmers

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available that take advantage of different chemical reactions to provide a gentle heat for an extended period of time. One type of hand warmer contains a mixture of dry chemicals, including among other things powdered iron and

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Alex Brinley, Charis
Conwell, and Siena
Joy for our AP
Chemistry class.

Designing a Hand
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Calculate the heat
released as the
solid Designing A
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Lab Answers Heat
approximately 125
mL of distilled

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water to 60—70 °C in a 250-mL beaker. Using heat-resistant gloves, measure 100.0 mL of the hot water in a 100-mL graduated cylinder. Measure and record the temperature of the hot water.

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Warmer Ap Lab Answers

AP Chemistry:
Designing an
Effective Hand
Warmer Designing
a Hand Warmer
Purpose: To
determine the best
solute to use to
make a safe hand
warmer. Procedure:
Data Part A: (table)
Part B and C:

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(table)
Calculations: Part
A: #13- calculate
calorimeter

constant Part B and
C: #7 Calculate the
molar heat of
solution q_{solution}
 $= -(q_{\text{aq}} + q_{\text{cal}}) =$
 $-(m \dots$

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Design and make a rechargeable handwarmer which, when it is needed, can be set off to provide a steady supply of heat at a comfortable temperature. ...

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race” competitions that help to put the fun back into chemistry, and gives students that “little bit extra” to motivate them ...

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investigate energy changes and calorimetry with formations of solutions. Students challenge themselves to design the best, all-around hand warmer.

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Applications. -

Lithium chloride,
the most effective
compound, was
also the most
dangerous and
expensive (6 times
more) to use. -

Lithium chloride
appears to be best

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choice for future
hand warmer
research and
design. - Lab helps
to illustrate the
idea that a balance
between safety,
cost and efficiency
must be struck in
engineering.

Designing a Hand
Warmer - ISU Lab
by Adam H

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How to Make A
Chemical Hand
Warmer In a small
zip-top bag, mix

1-1/2 tablespoons
iron filings, 1-1/1
tablespoons salt,
1-1/2 tablespoons
sand (or other...

Squeeze the air out
of the plastic bag
and seal it. It's a
good idea to place
the bag of chemical

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inside another bag,
remove the excess
air, and ...

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1. This lab demonstrated that various chemicals can have different energy changes.

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