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9.1 Introduction to Stoichiometry

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stoichiometry - mass mass problems*

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FACTORS AFFECTING RATE OF
REACTION *Stoichiometry: What is*

Stoichiometry? ~~Lesson 9.1~~ Line Plots

~~Stoichiometry: Converting Grams to
Grams~~ Chapter 9 Review part 2

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Mole, Grams to Grams, Mole Ratio

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CHAPTER 9 REVIEW Stoichiometry

MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? 40.07 g/mol

b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O

c. What is the mole ratio of O_2 to H_2O ?

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fewer steps are required to solve
stoichiometry problems when. ...

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Chapter 6: Chemical Bonding. 30 terms.

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Answer Key Microscopic: Two molecules

of hydrogen peroxide (in aqueous
solution) decompose to produce two

molecules of liquid water and one

molecule of oxygen gas. Chapter 9:

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Modern Chemistry 77 Stoichiometry

CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the

answer on the line to the left. Show all
your work in the space provided. 1. _____

The actual yield of a reaction is 22 g and
the theoretical yield is 25 g. Calculate the
percentage yield. 2. 6.0 mol of N₂ are
mixed with 12.0 mol of H

~~CHAPTER 9 REVIEW Stoichiometry~~

Stoichiometry b. Theoretically, how many
moles of NH₃ will be produced?

PROBLEMS Write the answer on the line
to the left, Show all your work in the space
provided. 1 88% The actual yield of a

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Answers
reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the ...

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Stoichiometry. SECTION 2. PROBLEMS
Write the answer on the line to the left.
Show all your work in the space provided.

1. The following equation represents a

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laboratory preparation for oxygen gas: ...

CHAPTER 9 REVIEW ...

~~CHAPTER 9 REVIEW~~

Chapter 9: Standard Review Worksheet 1.

Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

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Answers ...

Chapter 9 - Stoichiometry. 9-1

Introduction to Stoichiometry.

Composition Stoichiometry - deals with mass relationships of elements in

compounds Reaction Stoichiometry -

Involves mass relationships between reactants and products in a chemical

reaction. I. Reaction Stoichiometry

Problems A. Four problem Types, One Common Solution.

~~Chapter 9~~ Stoichiometry

Chapter 9 Review Stoichiometry Answers

CHAPTER 9 REVIEW Stoichiometry

MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following

equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g)$

+ $2H_2O(g)$ a. What is the value of the

coefficient x in this equation? 40.07 g/mol

b. What is the molar

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Section 2~~

CHAPTER 9 REVIEW Stoichiometry

MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a.

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