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Chapter 12.1, 12.2 Stoichiometry p1 *Unit 1 chapter 12 stoichiometry* Chapter 12 Stoich Limiting Reactant

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Chapter 12 Stoichiometry127. SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353 – 358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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Introduce the term sto- ichiometry in your own words. Stress that stoichiometry allows students to calculate the amounts of chemical sub- stances involved in chemical reactions using information obtained from bal- anced chemical equations.

12.1 The Arithmetic of Equations 12

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Chapter 3: Stoichiometry – Guided Reading Section 3.1 – 3.2 1. True or False? Most hydrogen atoms have a mass of 1.008 amu. Justify your answer. If true, explain why it is true. If false, what mass do most hydrogen atoms have? False, 1.008 amu is actually hydrogen ' s average mass, NO atom of hydrogen actually has the mass of 1.008 amu. 2.